The Mole [](http://www.google.co.uk/url?sa=i&rct=j&q=&esrc=s&source=images&cd=&cad=rja&uact=8&ved=0CAcQjRw&url=http%3A%2F%2Fwww.exmoormolecatcher.co.uk%2F&ei=hXaRVZ2mC8GV7Aa6xaCIDg&bvm=bv.96783405,d.ZGU&psig=AFQjCNE6eSsrhQKckYbwtbrn2K3eCu-oZw&ust=1435682798748078)

Read page 5 of your textbook.

What is a mole?...................................................................................

How many particles are there in a mole?

…………………………………………………………………….

What is a molar solution?

………………………………………………………………………

The molarity of a solution = number of moles of solute

number of litres of solution

Look at the worked example for making a molar solution of sodium chloride.

Do the same for a molar solution of:-

1. NaOH
2. MgCl2
3. HCl
4. CaCl2

Extension questions:

1. How many grams of NaOH would you need to make 450ml of a 0.250M NaOH solution? (answer – 4.5g)
2. How many grams of CaCl2 would you need to make 750ml of a 0.1M CaCl2 solution? (answer – 8.3g)